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## COURSE : Full Chemistry (chap 1-5)

### (1 mark each )

1. What products will be formed when ferrous sulphate crystals are heated ?
2. Write Black and white photography reaction ?
3. Define double displacement reaction ?
4. Litmus solution is a \_\_\_\_\_\_\_\_ dye .
5. What are olfactory indicators ?
6. Zn + NaOH ------ \_\_\_\_\_\_\_ + \_\_\_\_\_\_\_\_
7. CaCO3 + H20 + CO2 ------- \_\_\_\_\_\_\_\_\_\_
8. What is pH?
9. Write chemical formula of Milk of magnesia ?
10. What is the common name of compound CaOCl2?
11. Name the sodium compound which is used for soft\_\_\_\_\_\_\_\_ dye .
12. Define Neutralisation reaction ?
13. What are amphoteric oxides ? Give examples .
14. Name non metal which is liquid at room temperature ?
15. Name any non-metal which is lustrous ?
16. Name any metal which can be cut with knife?
17. Name two metals which donot react with neither hot nor cold water but only with steam?
18. Name metals which do not react with water at all ?
19. Name those metals which when react with HNO3 can evolve H2 gas ?
20. Define tetravalency?
21. Define catenation?
22. What is glacial acetic acid ?
23. Describe modern periodic law?
24. Describe mendeleev periodic law ?

## (2 marks each )

1. Nitrogen (atomic no. 7 ) and Phosphorus(atomic no. 15 ) belong to group 15 of the periodic table . Write the electronicconfiguration of these two elements . Which of these will be more electronegative ? why?
2. An atom has electronic configuration 2,8,7.
3. What is the atomic no. of this element .
4. To which of the following elements would it be chemically similar?
5. N(7)
6. F(9)
7. P(15)
8. Ar(18)
9. Which element has
10. Two shells, both of which are completetly filled with electrons
11. The electronic configuration 2,8,2 ?
12. A total of three shells , with four electrons in its valence shell?
13. Twice as many electrons in its second shell as in its first shell?
14. Draw the structures for the following compounds ?
15. Ethanoic acid
16. Bromopentane
17. Butanone
18. Hexanal
19. How would you distinguish experimentally between an alcohol and a carboxylic acid?
20. Write equation of both amphoteric oxides ?
21. Why sodium is immersed in kerosene ?
22. Write equations for the reaction of
23. Iron with steam
24. Calcium and potassium with water
25. (a) Why should curd and sour substances not be kept in brass and copper vessels? Which gas is usually liberated when an acid reacts with a metal? Illustrate with an example. How will you test for the presence of this gas?
26. Metal compound A reacts with dilute hydrochloric acid to produce effervescence. The gas evolved extinguishes a burning candle. Write a balanced chemical equation for the reaction if one of the compounds formed is calcium chloride.
27. Why does dry HCl gas not change the colour of the dry litmus paper?
28. Why is it advised not to add water to acid explain.
29. How is the concentration of hydronium ions (H3O+) affected when a solution of an acid is diluted? (b) How is the concentration of hydroxide ions (OH– ) affected when excess base is dissolved in a solution of sodium hydroxide?
30. Fresh milk has a pH of 6. How do you think the pH will change as it turns into curd? Explain your answer.
31. Identify the substances that are oxidised and the substances that are reduced in the following reactions.

(i) 4Na(s) + O2(g) → 2Na2O(s)

(ii) CuO(s) + H2(g) → Cu(s) + H2O(l)

## DO YOUR BEST AND LEAVE THE REST ON HIM

ALL THE BEST WITH YOUR REVISION